

A METHOD AND COMPOSITION FOR THE ACCELERATED IN VIVO REMOVAL OF ETHANOL

Field of the Invention

5 This invention relates to compositions and processes which accelerate the oxidation of alcohol in human blood. More particularly, the invention provides for the enhanced in vivo oxidation of ethanol and/or methanol.

Background of the Invention

10 Alcohol consumption and its consequences have long been a challenge to the statutes that govern our society. High levels of ethanol in the blood have far reaching consequences to the user and society, as indicated by 16000 highway deaths and 1.6 million drinking related arrests per year in the United States.

15 Therefore there is a need to provide a remedy which can be administered to accelerate removal of alcohol from an individual's bloodstream and therefore to shorten the period during which the individual is impaired.

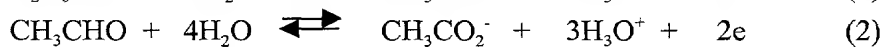
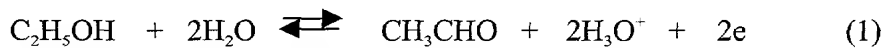
There is also a need to provide a remedy to an individual who is affected by alcohol to such a degree that medical intervention is required.

20 The source of oxygen, both mechanical and chemical (including the use of catalase in combination with hydrogen peroxide) is an essential part of the process, both these enzymes are NOT oxygen dependent but rather use NAD^+ as the oxidizing agent.

Even though the combination can include other enzymes, which regenerate NADH to NAD^+ , using glycerol dehydrogenase/dihydroxyacetone, the presence of a liberal supply of oxygen is an essential part of the method.

Summary of the invention

25 Formally, the oxidation of ethanol to acetaldehyde and then to acetic acid can be generically represented by the reactions:



It is therefore the object of this invention to accelerate the rates of the forward reactions in the equilibria 1 and 2, by introducing additives which catalyze the forward reactions or which remove product from the right hand side of the equilibria 1 and 2 to favor their shifting to the right.

Various additives provide a remedy through an effective and rapid reduction of the alcohol level in blood of the consumer at a reasonably short time, in essence a sobriety agent. Such remedy can take the form of a liquid, a pill, a capsule, a patch or a nasal spray or any other method known to medical science and any combination thereof. Such remedy can be self-administered or, in the same or different formulation be administered at a health care facility where it can be applied intravenously, intraperitoneally or in any other method known in medical practice. The product shall apply to all alcohol containing beverages including but not limited to all beers, wines and hard liquors now in use.

Brief Description of the Drawings

Figure 1a is a graph of the oxidation of NADH by vanadyl sulfate at pH 9.25.

Figure 1b is a graph of the oxidation of NADH by sodium phosphotungstate at pH 9.25.

Figure 1c is a graph of the oxidation of NADH by potassium ferricyanide at pH 9.25.

Figure 1d is a graph of the oxidation of NADH by ammonium molybdate at pH 9.25.

Figure 2a is a graph of the oxidation of NADH by vanadyl sulfate at pH 7.25.

Figure 2b is a graph of the oxidation of NADH by sodium phosphotungstate at pH 7.25.

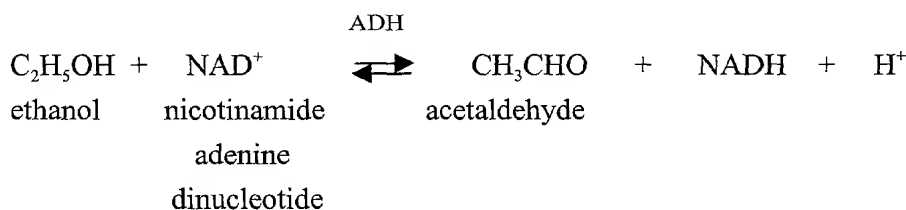
Figure 2c is a graph of the oxidation of NADH by potassium ferricyanide at pH 7.25.

Figure 2d is a graph of the oxidation of NADH by ammonium molybdate at pH 9.25.

Detailed Description of the Preferred Embodiments

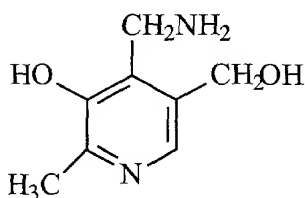
The present invention relates to compositions which when administered to a person enhance the in vivo physiological process of removing oxidizing alcohol and particularly from the blood. In this disclosure, the term "alcohol" when used without qualification is understood to refer to ethanol. However, certain

embodiments of the invention as noted will also be applicable to methanol. As alcohol enters the blood stream, the liver begins the process of detoxification through the production of the enzyme Alcohol Dehydrogenase, ADH, which catalyzes the oxidation by Nicotinamide Adenine Dinucleotide, NAD⁺, of ethanol to acetaldehyde and Dihyronicotinamide Adenine Dinucleotide, NADH, as shown.

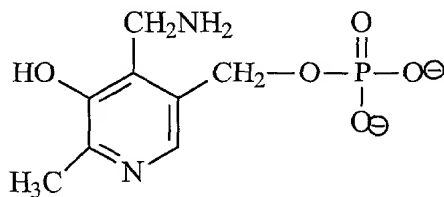


The equilibrium constant of this reaction is 1, indicating that it is rather inefficient, allowing sufficient amounts of alcohol to remain in the blood and to cross the blood-brain barrier. It is the purpose of this invention to find ways of increasing the efficacy or accelerating this reaction.

In one embodiment of this invention, accelerated oxidation is accomplished by introducing a chemical accelerator of the ADH enzyme, driving the reaction toward acetaldehyde. Such accelerators include zinc ions (as ADH is a Zinc Enzyme), pyridoxamine or pyridoxamine phosphate (formulae 1a and 1b) or a combination thereof.



1a

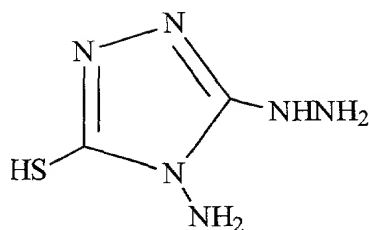
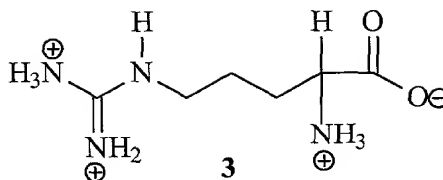
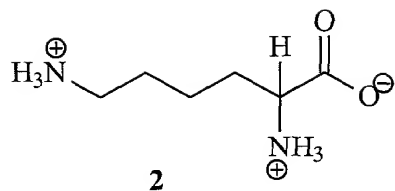


1b

It is a particular advantage in the case of one of the above accelerators (zinc) that it is an essential element or a part of the Vitamin B6 system (pyridoxamine).

In another embodiment of this invention, the reversible reaction above can be shifted toward irreversibility by the removal of one or more of the products (Le Chatelier's Principle). This is accomplished through the introduction of a reactive agent that by reacting with acetaldehyde (removing it from the equilibrium) drives the reaction forward. Such an agent includes pyridoxamine or pyridoxamine phosphate (formulae 1a and 1b) or the basic amino acids lysine (formula 2) and arginine (formula 3). A solubilized form of "Purpald" (formula 4) may also be useful. Other agents reactive with acetaldehyde include thiamine.

A sufficient quantity of the reactive agent is administered to provide an *in vivo* concentration of the reactive agent at least chemically equivalent to an amount of acetaldehyde resulting from the oxidation.



Lysine and arginine being essential amino acids, their introduction would not cause an undue risk to the user. Further, being basic, these amino acids can help neutralize the acid formed during the oxidation, thus further shifting the equilibrium toward product.

In the presence of a base, the acetic acid produced by equilibrium (2) is neutralized, which pulls the reaction to the right. This consumes acetaldehyde from equilibrium (1), which is thus pulled to the right and consumes more ethanol.

Any base can be used provided that it is not harmful to the body. For example, the base can be sodium carbonate, sodium bicarbonate, trisodium phosphate, disodium hydrogen phosphate or tris(hydroxymethyl)-aminomethane. A sufficient quantity of the base is administered to provide an *in vivo* concentration of the base at least chemically equivalent to the acid resulting from the oxidation of the ethanol.

The removal of acetaldehyde can also be accomplished by its enzymatic oxidation to acetic acid as shown.



In one embodiment of the invention this can be accomplished by adding acetaldehyde dehydrogenase or an accelerator thereof or a combination of the two.

When *in vivo*, the dehydrogenase should be present in an amount at least chemically equivalent to the amount of blood alcohol or, more precisely, to acetaldehyde, the oxidation product of alcohol. This translates to a concentration in the range 0.1 – 10 I.U./L

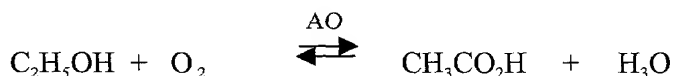
Optionally the dehydrogenase may be stabilized with zinc ions since the dehydrogenase might otherwise degrade. Preferably, the concentration of zinc ions is at least 1% of the molar concentration of the dehydrogenase.

By increasing the amount of one of the reactants, the equilibrium can be shifted toward product. Thus, a formulation containing NAD^+ can be useful. A sufficient quantity of NAD^+ is administered to provide an *in vivo* concentration of NAD^+ in the range 0.05% to 5% of the maximum expected *in vivo* molar concentration of ethanol.

Any combination of the above may also be beneficial. Thus, a formulation containing acetaldehyde dehydrogenase, NAD^+ , pyridoxamine phosphate and zinc arginate or lysinate (or both) can be useful.

In another embodiment of the invention, the acetaldehyde can be chemically or enzymatically converted to an innocuous derivative, which can later be removed from the body. These derivatives include, but are not limited to, esters, ethers, acetals, ketals and urethanes.

Under proper aerobic conditions and with sufficient acid neutralizing agents, the aerobic oxidation of alcohol to acetic acid catalyzed by Alcohol Oxidase, AO, as shown,



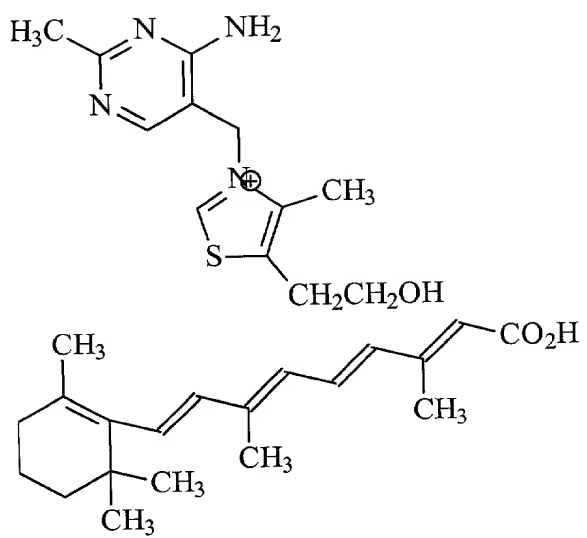
may also be useful, particularly in combination with any of the above schemes.

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In another embodiment of the invention, the acetic acid formed in either the acetaldehyde dehydrogenase or the alcohol oxidase reaction, or the combination of the two, can be converted to methane and carbon dioxide through a methanogenetic enzyme, thus driving the alcohol oxidation further toward
5 irreversibility, as shown in equation 4.

In another embodiment of this invention, the enzymatic reaction can be accelerated by a charge-transfer agent such as thiamine, (formula 5), retinoic acid (formula 6), an isoflavonoid or a pyranoside thereof (e.g. daidzein (formula 7) or its 7-glucoside, commonly known as aloin, (formula 8)), and 4,5-Dihydro-4,5-
10 dioxo-1H-pyrrolo[2,3-f]quinoline-2,7,9-tricarboxylic acid, also known as either pyrroloquinolinequinone (PQQ) or methoxatin (formula 9), or any combination thereof. lipoic acid, retinoic acid, retinal, retinol, and derivatives and analogs thereof. A sufficient quantity of the charge-transfer agent is administered to provide an *in vivo* concentration of the charge-transfer agent in the range from
15 0.1% and 2% of the maximum expected *in vivo* molar concentration of ethanol.

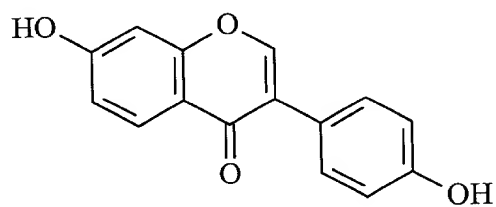
In a further embodiment of the invention, various accelerants can be used to supply energy to the forward reactions of equilibria 5 and 6. Such accelerants include adenosine 5'-triphosphate, adenine-9- β -D-arabinofuranside 5'-triphosphate, 2'-deoxyadenosine 5'-triphosphate, and 2',3'-dideoxyadenosine 5'-
20 triphosphate. They also include carbohydrates such as fructose, arabinose, ribose, deoxyribose, and their phosphorylated derivatives. A sufficient quantity of the accelerant is administered to provide an *in vivo* concentration in the range from 1% to 100% of the maximum expected *in vivo* molar concentration of ethanol.

25 In another embodiment of this invention, a combination of some or all of the above with substances that prevent or delay the absorption of gastric alcohol into the blood is also envisaged. Such substances include unsaturated fatty acids, dietary fiber, and surfactants such as oleic acid, lecithin, the plant surfactant saponin and taurine (formula 10).

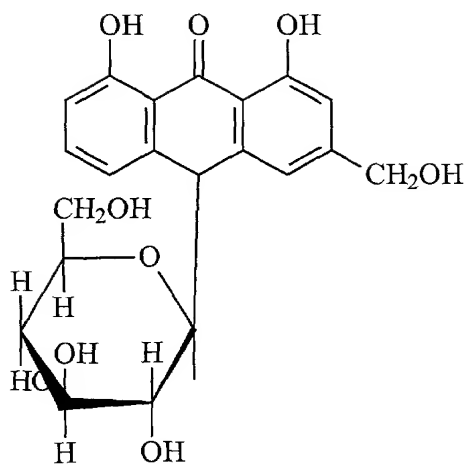


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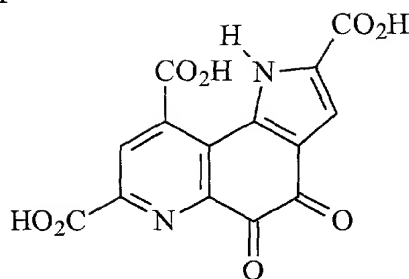
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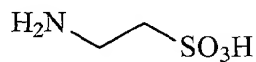
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8



9



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The concentration of the surfactant in the composition should be in the range 0.02% and 0.2% by volume.

Some surfactants which also act as charge transfer agents can be used.

These include lipoic acid, retinoic acid, retinal, retinol, and derivatives and analogs thereof. The concentration of the combined surfactant and charge-transfer agent should be between 0.1% and 2% of the maximum molar concentration of ethanol. A sufficient quantity of the combined surfactant and charge-transfer agent is administered to provide an *in vivo* concentration of the surfactant and charge-transfer agent between 0.1% and 2% of the maximum expected molar concentration of ethanol.

In yet another embodiment of this invention, the acceleration is effected by multivalent transition metal ions and their derivatives capable of acting as a pseudo enzyme in a dehydrogenase reaction. The ions include, but are not limited to, ions of elements of Groups IVa through VIII of the Periodic Table and their complexes or any combination thereof.

Preferred sources of such ions include: vanadyl sulfate; potassium ferricyanide; ammonium iron (III) citrate; ammonium molybdate; ammonium phosphomolybdate; sodium tungstate; and sodium phosphotungstate. Other sources include ammonium manganese (III) sulfate; zirconium (IV) EDTA; niobium (IV) EDTA; tetrakis(tropolinato) niobium (V) chloride; tetrakis(tropolinato) tantalum (V) chloride; cobalt (III) hexammine chloride; and chromium (III) picolinate. A sufficient amount of transition metal ion administered to provide an *in vivo* concentration of the ion in the range 0.05% to 2% of the maximum expected *in vivo* molar concentration of ethanol.

In the foregoing disclosure some reactions, such as those involving the enzyme ADH, are specific to ethanol, while others may apply also to methanol. In general, the multivalent transition metal ions and their complexes mentioned above would be expected to oxidize methanol and ethanol alike, via reactions which may sometimes involve NAD^+ . It has been found in particular that vanadyl sulfate/ NAD^+ is effective in oxidizing methanol. This invention therefore encompasses the *in vivo* oxidation of both ethanol and methanol in the blood.

Optionally, any of the embodiments can also include dietary constituents such as dietary fiber, garlic oil and onion oil, or medications including such commonly available pain relieving ingredients as aspirin, ibuprofen and acetomenaphin.

Examples

Examples 1 - 40 were designed to demonstrate, in a semi-quantitative way the effectiveness of the claimed invention.

In each of the examples, aliquots of reagents were mixed and the final solution allowed to stand for a specified time, after which it was treated with an

excess of a reagent containing 2,4-dinitrophenylhydrazine (DNP) in a strongly acidic solution. The reagent quenches any oxidative reaction of alcohol as well as detects the formation of acetaldehyde. An orange, crystalline precipitate indicates a positive reaction. The DNP reagent is a fresh solution of 0.5 gm. of 2,4-dinitrophenylhydrazine in 20mL methanol and 5mL concentrated sulfuric acid, diluted to 50 mL of deionized distilled water.

Example 1

To a 0.05M solution of ethanol in a 1M (pH 9.25) phosphate buffer (10 μ L) are added 10 μ L each of a solution 0.08M in NAD^+ (in deionized distilled water) and 4U/mL Alcohol Dehydrogenase (from Equine liver) in the above buffer, and diluted to 100 μ L with buffer. At the end of a specified time, it was quenched with the DNP reagent. A positive reaction with DNP was obtained when the alcohol solution was allowed to stand for 2 minutes.

Example 2

The procedure in Example 1 was repeated with the omission of the enzyme solution. The solution tested negative even after being allowed to stand for 15 minutes, indicating that NAD^+ cannot oxidize alcohol without a catalyst.

Example 3

The procedure in Example 1 was repeated with the addition of 10 μ L of a 0.01M vanadyl sulfate solution in deionized distilled water. The solution tested positive after being allowed to stand for 1 minute, indicating some acceleration.

Example 4

The procedure in Example 3 was repeated with the omission of the enzyme solution. The solution tested positive after being allowed to stand for 2 minutes, indicating that vanadyl sulfate can act as a catalyst or a "pseudo" enzyme for the oxidative reaction.

Example 5

The procedure in Example 4 was repeated with the omission of NAD^+ . The solution tested negative even after being allowed to stand for 15 minutes, indicating that vanadyl sulfate cannot oxidize alcohol in the absence of NAD^+ .

Example 6-10

Examples 1-5 were repeated, with almost identical results, using sodium phosphotungstate in place of vanadyl sulfate.

Example 11-15

5 Examples 1-5 were repeated, with almost identical results, using ammonium molybdate in place of vanadyl sulfate.

Example 16-20

Examples 1-5 were repeated, with almost identical results, using potassium ferricyanide in place of vanadyl sulfate.

10 Example 21-25

Examples 1-5 were repeated, with almost identical results, using ammonium ferric citrate in place of vanadyl sulfate.

Example 26

15 The procedure in Example 2 was repeated using a pH 7.25 1M phosphate buffer. The solution tested negative even after being allowed to stand for 15 minutes, indicating that the enzyme is ineffective *in vitro* at pH 7.25

Example 27

20 The procedure in Example 2 was repeated using a pH 7.25 1M phosphate buffer. The solution tested negative even after being allowed to stand for 15 minutes, indicating that NAD^+ cannot oxidize alcohol without a catalyst.

Example 28

25 The procedure in Example 3 was repeated using a pH 7.25 1M phosphate buffer. The solution tested positive after being allowed to stand for 2 minutes, indicating that vanadyl sulfate can act as an accelerator for the oxidative reaction, even at the lower pH 7.25.

Example 29

30 The procedure in Example 4 is repeated using a pH 7.25 1M phosphate buffer. The solution tested positive after being allowed to stand for 2 minutes, indicating that vanadyl sulfate can act as a catalyst or a "pseudo" enzyme for the oxidative reaction, even at the lower pH 7.25.

Example 30

The procedure in Example 5 is repeated using a pH 7.25 1M phosphate buffer. The solution tested negative even after being allowed to stand for 15 minutes, indicating that vanadyl sulfate cannot oxidize alcohol in the absence of NAD⁺.

Example 31-35

Examples 25-30 were repeated, with almost identical results, using sodium phosphotungstate in place of vanadyl sulfate.

Example 36-40

Examples 25-30 were repeated, with almost identical results, using ammonium ferric citrate in place of vanadyl sulfate.

Examples 41 and 42 were designed to demonstrate together, in a semi-quantitative way, the effectiveness of a particular transition metal (vanadium) compound in catalyzing the oxidation of acetaldehyde to acetic acid.

Example 41

To a 0.0002 M solution of acetaldehyde in deionized distilled water (1 μ L) are added 10 μ L of a solution 0.0004 M in NAD⁺ (in deionized distilled water) and 1U/mL Acetaldehyde Dehydrogenase (from Baker's Yeast) in a 1M (pH 9.25) phosphate buffer (1 μ L), and diluted to 100 μ L with the above buffer. At the end of a specified time, it was quenched with the DNP reagent. A positive reaction with DNP was obtained when the solution was allowed to stand for 2 minutes; an ambiguous reaction after 5 and 10 minutes, and a negative reaction after 20 minutes. This indicates that the most of the acetaldehyde had been converted.

Example 42

The procedure in Example 41 was repeated replacing the Acetaldehyde Dehydrogenase was replaced by a solution 0.004 M in vanadyl sulfate (in deionized distilled water) (1 μ L). An ambiguous reaction was obtained with the DNP reagent after 2 and 5 minutes; a negative reaction was obtained after 10 minutes. This indicates that vanadyl sulfate is at least as effective as Acetaldehyde Dehydrogenase in catalyzing the oxidation of acetaldehyde.

Examples 43-50 were designed to illustrate that various transition metal compounds can effect the oxidation of NADH, thus recycling NAD^+ , a major feature of any catalyst.

Example 43

A 0.01M solution of NADH in a 1M (pH 9.25) phosphate buffer (10 μL) is diluted to 1900 μL with the buffer and mixed with a 0.01M solution of vanadyl sulfate (10 μL) in a multi-cavity plate. The rate of oxidation of NADH is followed over time as the decrease in the absorption at 340 nm. The results, as summarized in Table 1 and Figure 1 indicate the effectiveness of vanadyl sulfate and potassium ferricyanide, and to a lesser extent, ammonium molybdate and sodium phosphotungstate, the latter two after an induction period.

Example 44

The procedure in Example 43 is repeated replacing vanadyl sulfate with potassium ferricyanide.

Example 45

The procedure in Example 43 is repeated replacing vanadyl sulfate with sodium molybdate.

Example 46

The procedure in Example 43 is repeated replacing vanadyl sulfate with sodium phosphotungstate.

Examples 47-50

The procedures in Examples 43-46 are repeated using a 1M (pH 7.25) phosphate buffer solution. The results, as summarized in Table 2 and Figure 2, indicate the effectiveness of vanadyl sulfate and potassium ferricyanide, and to a lesser extent sodium molybdate, the latter after an induction period.

Table 1

Kinetics of Transition Metal Oxidation of NADH at pH 9.25

Additive	None	Vanadyl Sulfate	Phospho-tungstate	Molybdate	Ferricynide
Slope (OD/Min)	0	.0112	.007*	.0035*	.0432
Slope (Mole/L/Min)	0	.000005	.0000006	.0000003	.00002

Table 2

Kinetics of Transition Metal Oxidation of NADH at pH 7.25

Additive	None	Vanadyl Sulfate	Phospho-tungstate	Molybdate	Ferricynide
Slope (OD/Min)	0	.0079	0	.0025*	.0285
Slope (Mole/L/Min)	0	.0000036	0	.00000012	.000013

In Examples 51-53, the rate of ADH catalyzed NAD⁺ oxidation of alcohol is followed by the rise in absorption at 340 nm, with and without additives. When a solvent (methanol) is needed to dissolve the additive, an equal amount is added to the "blank" or "control". The results, as summarized in Table 3, indicate the effectiveness of Daidezin, Aloin and Methoxatin as accelerators.

Example 51

To a 0.01M solution of ethanol (10μL) in 1M pH 9.25 phosphate buffer, in a multi-well plate, is added a solution of aloin, of the indicated concentration (10μL) and 10μL of a solution containing 0.2U/mL ADH and diluted to 180μL with buffer. To the mixture is

added 20 μ L of a 0.01M solution of NAD⁺ in deionized distilled water, and the absorbance at 340 nm is followed with time.

Example 52

The procedure in Example 51 is repeated using Daidezin.

Example 53

The procedure in Example 51 is repeated using Methoxatin.

Table 3
Relative rates of Ethanol Oxidation in presence of Additives
(Control = 1)

Aloin (.000005M)	Aloin (.00001M)	Daidezin (.000005M)	Daidezin (.00001M)	Methoxatin (.000005M)	Methoxatin (.00001M)
1.07	1.19	1.28	1.38	1.29	1.59

While the invention has been described in connection with various embodiments, it is not intended to limit the scope of the invention to the particular form set forth, but on the contrary, it is intended to cover such alternatives, modifications, and equivalents as may be included within the spirit and scope of the invention as defined by the appended claims.